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Red Hat Enterprise Linux 7
System Administrator's Guide

Deployment, Configuration, and Administration of Red Hat Enterprise Linux
7

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Answer Key

Chapter 14

Practice Problems

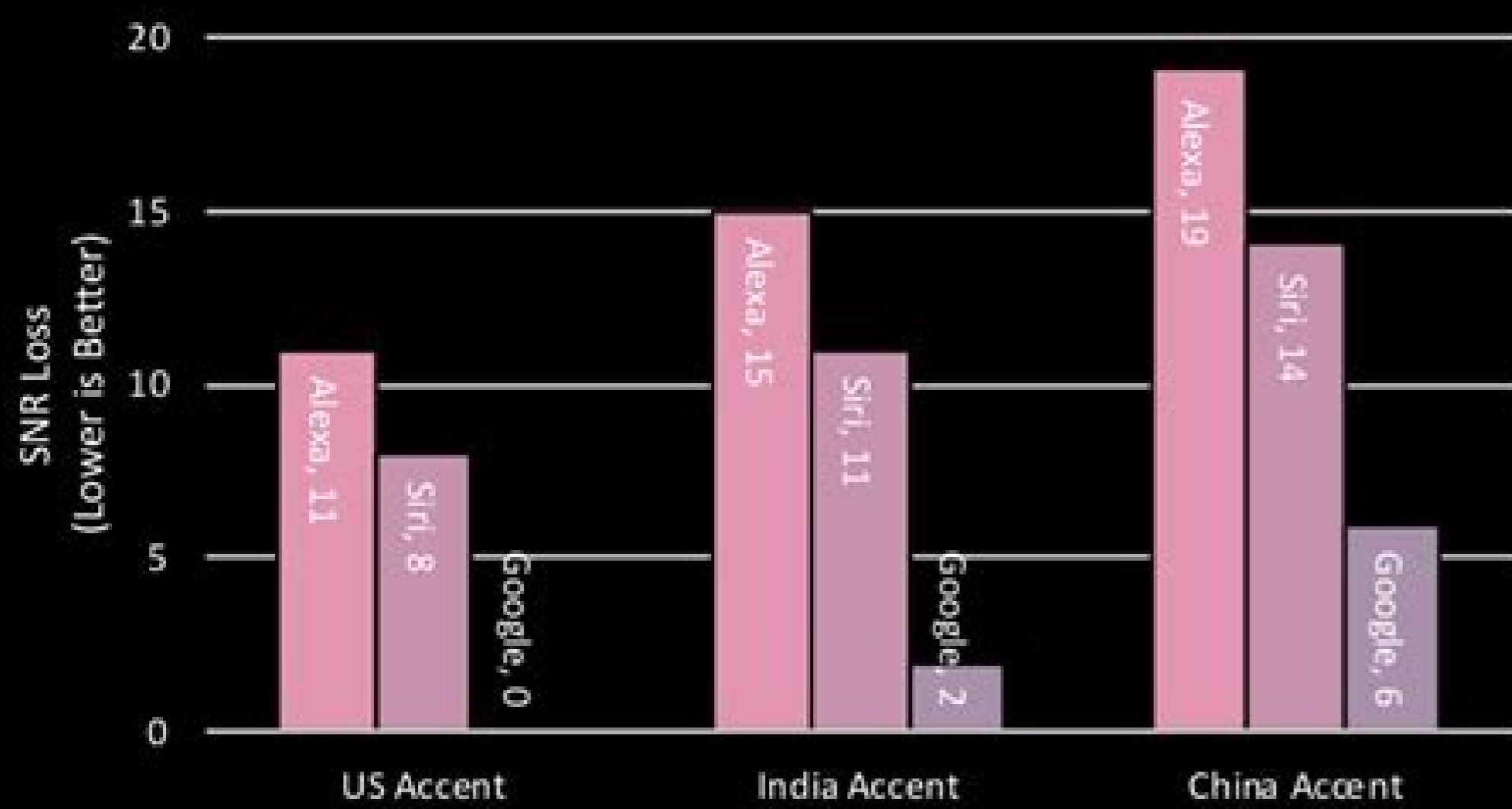
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|--|--|
| 1. a. 61 g CaCl ₂ /100 g H ₂ O | 12. 0.802M |
| b. 90 g CaCl ₂ /100 g H ₂ O | 13. 0.350m |
| c. 5 g Ce ₂ (SO ₄) ₃ | 14. 0.0300m |
| d. 1 g Ce ₂ (SO ₄) ₃ | 15. a. water, 0.830; ethanol, 0.170 |
| e. 60 g KClO ₃ | b. water, 0.9853; calcium chloride, 0.0147 |
| 2. 94 atm | c. carbon tetrachloride, 0.644; benzene, 0.356 |
| 3. 2.1 atm | 16. BP = 102.06°C; FP = -7.48°C |
| 4. 0.852% | 17. 2.33m |
| 5. 7.35 g | 18. a. suspension |
| 6. 492.65 g | b. solution or colloid |
| 7. 26% | c. colloid |
| 8. 0.868 L | d. solution |
| 9. 0.328M | |
| 10. 1.09M | |
| 11. 326 mL | |

Chapter 14 Review

19. Solvation is a process in which solvent particles surround solute particles to form a solution. A substance dissolves when attractive forces between solvent and solute particles are strong enough to overcome the attractive forces holding the solute particles together.
20. A saturated solution contains the maximum amount of dissolved solute for a given amount of solvent at a specific temperature and pressure. An unsaturated solution contains less dissolved solute for a given temperature and pressure than does a saturated solution. A supersaturated solution contains more dissolved solute than does a saturated solution at the same temperature.

Chemistry: Matter and Change 1 Solving Problems: A Chemistry Handbook
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